Periodic Table of Schroeder Valence (e-) Electron Circle Vacancies

H He Li Be B C N O F Ne Na Mg Al 3+ Si 4+ P S Cl Ar K Ca Sc Ti V Cr Mn Fe Co Ni Cu Zn Ga Ge As Se Br Kr Rb Sr Y Zr Nb Mo Tc Ru Rh Pd Ag Cd In Sn Sb Te I Xe Cs Ba La Hf Ta W Re Os Ir Pt Au Hg Tl Pb Bi Po At Rn Fr Ra Ac Th Pa U

Ce Pr Nd Pm Sm Eu Gd Tb Dy Ho Er Tm Yb Lu

Cesium Fluoride is simply filling the one empty Schroeder e-Circle Vacancy

CsF (solid) Cs + F* = Cs F

Considered to be the strongest ionic bond...
see how F is right hand TOP ('ElectroNEGATIVE')
and Cs is lower LEFT ('ElectroPOSITIVE')

F Ca F

Just fill in the Blanks" for Calcium Fluoride....

CaF_2

F F C F F Cl

...or a ChloroFluoroCARBON

CF_3Cl

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